

Chemistry



9th Class

Sindh Board Notes

Chapter # 10

Chemical Energetics



Fill In Blanks

پنجاب، سندھ، بلوچستان، خیبر پختونخواہ، بورڈز کے تمام نوٹس سابقہ پیپرز، اس ویب سائٹ سے فری ڈاؤن لوڈ کریں۔

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→ CHEMICAL ENERGETICS

1. The heat given out in a chemical reaction is called exothermic reaction.
2. In endothermic reaction, heat is taken in.
3. Heat evolved or absorbed during chemical reaction at constant pressure is called enthalpy of reaction.
4. Acid base reaction is called neutralization reaction.
5. The chemical reactions during which materials changes are associated with change in heat energy are called thermo-chemical reactions.
6. The branch of chemistry which deals with the study of heat changes in chemical reactions is called thermo chemistry.
7. There are two types of thermo-chemical reactions i.e. exothermic and endothermic chemical reactions.
8. In Greek EXO means out of or to evolve and THERM means heat.
9. An exothermic reaction is the chemical change during which heat is given out or released.
10. The change of heat is represented by ΔH and it is shown by negative sign when heat is evolved.
11. 286 KJ/mole of heat energy is released when 1 mole of H₂ reacts with $\frac{1}{2}$ mole of O₂ to form 1 mole of H₂O.
12. In exothermic reactions, heat flows from the system to surroundings and container becomes hot.

13. In **exothermic** reaction, total energy of the reactants is greater than total energy of products.
14. The flameless radiation heater contains a mixture of **Mg, Fe** and **NaCl**.
15. ENDO means to **absorb** and THERM means **heat**.
16. Endothermic reaction is the chemical change during which heat is **absorbed or taken in**.
17. The sign of ΔH is positive when the reaction is **endothermic**.
18. During the endothermic reaction heat flows from the **surroundings** to the **system** and container becomes **cold**.
19. In exothermic reaction total energy of the **products** is greater than the total energy of **reactants**.
20. The energy possessed by a substance is called **heat contents** of that substance.
21. The heat evolved or absorbed at constant pressure is called as **enthalpy** of the reaction.
22. The enthalpy of a substance is represented by "**H**".
23. Δ Signifies the change in **property**.
24. If the enthalpy of the products is greater than the enthalpy of the reactants, then the sign of ΔH will be **positive (+ve)** and over all reactions is **endothermic**.
25. If the enthalpy of product is smaller than the reactants, then the sign of ΔH will be **negative (-ve)** and overall reaction is **exothermic**.
26. The heat absorbed or evolved during thermo-chemical reaction is called **heat of reaction**.
27. Exothermic and endothermic reaction can be easily be detected by touching the **vessel** before and after the **chemical reaction**.
28. The increase in temperature indicates that reaction is **exothermic**.
29. The decrease in temperature indicates that the reaction is **endothermic**.

30. The accurate value of ΔH can be determined by using calorimeter.
31. The reaction between an acid and base to form salt and water is called neutralization reaction.
32. Neutralization reaction is an example of exothermic reaction.
33. The heat of neutralization for any strong acid with strong base is approximately same.

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9th Class

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