

# Chemistry



9th Class

Sindh Board Notes

Chapter # 2

Chemical Combinations



Solved MCQs

پنجاب، سندھ، بلوچستان، خیبر پختونخوا، بورڈز کے تمام نوٹس سابقہ پیپرز، اس ویب سائٹ سے فری ڈاؤن لوڈ کریں۔

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## → CHEMICAL COMBINATIONS

- Mass can neither be created nor destroyed during a chemical reaction, is the law of**
  - ✓ conservation of mass
  - constant composition
  - multiple proportion
  - reciprocal proportion
- Law of Multiple proportion was proposed by**
  - Louise Proust
  - Lavoisier
  - ✓ John Dalton
  - Ritcher
- Berzelius performed the experiments on the preparation of**
  - H<sub>2</sub>S
  - ✓ PbS
  - ZnS
  - FeS
- A combination of symbols of atoms or ions, that are held together chemically in a compound is called**
  - empirical formula
  - ✓ formula
  - molecular formula
  - simplest formula
- 2 moles of CaCO<sub>3</sub> is equal to**
  - 100g
  - ✓ 200g
  - 88g
  - 160g
- 23 a.m.u. of Na is equal to**
  - ✓ 1 mole
  - 2 moles
  - 3 moles
  - 4 moles
- CaCO<sub>3</sub>  $\xrightarrow{\text{Heat}}$  CaO + CO<sub>2</sub> is an example of;**
  - addition reaction
  - combustion reaction
  - single displacement
  - ✓ decomposition reaction
- A reaction in which substances reacts with oxygen is**
  - decomposition reaction
  - single displacement reaction
  - addition reaction
  - ✓ combustion reaction
- Which type of reactions are neutralization and hydrolysis reaction?**
  - decomposition
  - single displacement
  - addition
  - ✓ double displacement



10. The number in front of the formula in a chemical equation is called
- expression
  - subscript
  - ✓ co-efficient
  - none of these.
11. Symbol  $\Delta$  represents
- state of matter
  - ✓ heat
  - light
  - catalyst
12. How many grams atoms are therein 5 moles of carbons?
- 12g
  - 24g
  - 36g
  - ✓ 60g
13.  $N_A$  represents
- atomic number
  - mass number
  - ✓ Avogadro's Number
  - mole
14. A reaction in which two or more substances react to form a single substance is called
- decomposition reaction
  - combustion reaction
  - displacement reaction
  - ✓ addition reaction
15. Avogadro's Number shows the number of
- ✓ particles
  - neutrons
  - protons
  - electrons
16. The formulas for all atoms in an empirical formula of a compound is
- Molecular formulas
  - Chemical formulas
  - ✓ Empirical formulas
  - Structural formulas
17. In 1961, the standard for atomic masses was designated by the isotope of:
- ✓ C-12
  - C-13
  - C-14
  - H-1
18. One mole of any substance contains
- Equal number of particles
  - Unequal number of particles
  - ✓ Avogadro numbers of particles
  - Dalton numbers of particles
19. Avogadro number is constant. Its value is:
- $3.02 \times 10^{22}$
  - ✓  $3.02 \times 10^{23}$
  - $3.02 \times 10^{21}$
  - $3.02 \times 10^{24}$
20. Benzene has molecular formula  $C_6H_6$ . Its empirical formula will be
- $C_{12}H_{12}$
  - $C_{18}H_{18}$
  - $C_2H_2$
  - ✓ CH

21. One mole of  $\text{CO}_2$  contains  $\text{CO}_2$   
a. ✓ 44g  
c. 32g  
b. 12g  
d. 22g
22. In water, the ratio between the masses of hydrogen and oxygen is:  
a. 2:16  
c. 2:1  
b. ✓ 1:8  
d. 1:2
23. The mass of one mole of substance is called:  
a. Formula mass  
c. Molecular mass  
b. Atomic mass  
d. ✓ Molar mass
24. In a chemical equation, reactants are written on the left hand side and products are written on the:  
a. ✓ Right  
c. Middle  
b. Left  
d. Upper
25. The symbol  $\rightleftharpoons$  represents a  
a. One way reaction  
c. ✓ Two way reversible reaction  
b. Half way reaction  
d. No way reaction

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FOR MORE NOTES, MCQS, ONLINE QUIZZES



# All Classes Chapter Wise Notes

Punjab Boards | Sindh Boards | KPK Boards | Balochistan Boards

AJK Boards | Federal Boards

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## 9th Class

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