

# Chemistry



9th Class

Sindh Board Notes

Chapter # 7

Solution And Suspension



Fill In Blanks

پنجاب، سندھ، بلوچستان، خیبر پختونخواہ، بورڈز کے تمام نوٹس سابقہ پیپرز، اس ویب سائٹ سے فری ڈاؤن لوڈ کریں۔

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## ➔ SOLUTION & SUSPENSION

1. Solution is a homogenous mixture of two or more substances.
2. The most common solvent in nature is water.
3. An aqueous solution is the solution when the liquid solvent is water.
4. 10%  $\frac{M}{M}$  solution contains 10 gram solute, dissolved in 90 g solvent.
5. M is the symbol for the concentration unit of molarity.
6. A solution is a homogenous mixture of two or more substances.
7. The component of solution present in smaller amount is called solute.
8. The component present in greater quantity is called solvent.
9. When water is solvent solution is called aqueous solution.
10. The states of matter solid, liquid and gas produce nine types of solutions on mixing.
11. Soft drinks are the examples of gas in liquid.
12. Cloud and steam are examples of liquid in gas.
13. Smoke is solid in gas.
14. Milk in tea is liquid in liquid.
15. The solubility of sugar in water at 0°C is 179 g/100 ml.
16. The solubility of gases in a liquid decreases with the increase in temperature.
17. The solubility of solids and liquids are not affected by pressure.
18. The solubility of a gas in a liquid is directly proportional to the pressure gas. This is called Henry's Law.

19. An unsaturated solution is that one which contains **less solute** than it has the capacity to **dissolve**.
20. A **saturated** solution is that which contains **maximum** amount of solute in a given solvent at room temperature.
21. The amount in grams of a solute required to be dissolved in 100 grams or 100 ml of a solvent to prepare a saturated solution at room temperature is called **solubility** of that solute.
22. A super saturated solution is that which contains **greater** amount of dissolved solute than that present in a saturated solution at the given temperature.
23. The process in which dissolved solute comes out of the solution and forms crystals is called **crystallization**.
24. **Crystals** are the homogenous solids, having regular and definite geometrical shape.
25. For crystallization super **saturated solution** is prepared.
26. Crystals of copper sulphate are called **blue vitriol**.
27. The formula for copper sulphate crystals is  **$\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$** .
28. In a solution the amount of solute dissolved in a given quantity of solvent is known as **concentration**.
29. Molarity, Molality, mole fraction, percentage and normality are the common methods of expressing the **strength** of a solution.
30. Molarity is defined as the number of **mole** of solute dissolved in **1** litre ( $1\text{dm}^3$ ) of a solution. Molarity is denoted by **M**.
31. The Molarity of any solution is found by dividing the number of moles of solute by the number of **litres** of solution.
32. 1 mole of NaOH (40g) dissolved in 1 litre of solution is said to be **1 mole (1M)** solution.
33. Number of moles of solute dissolved per 1000 (1kg) of solvent is called **molality**.
34. Molality is denoted by **m**.
35. 1 mole of  $\text{NaCO}_3$  dissolved in 100 grams of solvent is said to be **1 molal (1m)** solution.

36. **Mole fraction** is defined as the ratio of moles of solute or solvent to the total moles of the solution.
37. The sum of mole fractions of all components of a solution must be equal to **1**.
38. In **suspension** the solute particles are not soluble in solvent.
39. Suspension is **heterogeneous** mixture.
40. **Smoke** is a suspension of Carbon particles in a gas or air.
41. Mud is a **Suspension** of fine particles of solid in a small quantity of liquid.
42. **Foam (froth)** is a suspension of fine particles of a gas in a **liquid**.
43. Emulsion is a suspension of droplets of one liquid into another in which it is not **soluble**.

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