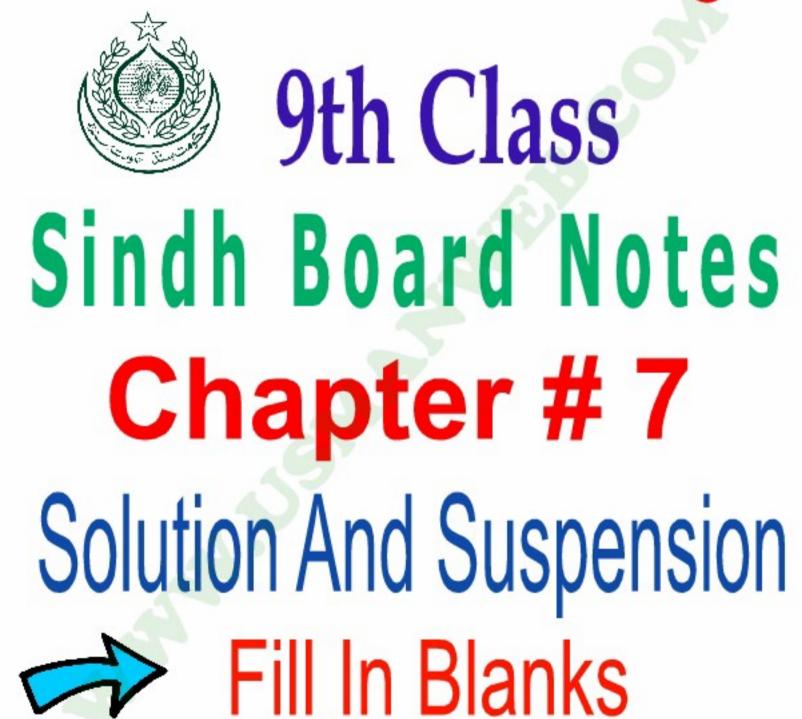
# Chemistry



پنجاب، سندھ، بلوچستان، خیبر پختونخواہ، بور ڈزکے تمام نوٹس سابقہ بیپرز،اس ویب سائیٹ سے فری ڈاو ملوڈ کریں۔

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Chapter # 07

Chemistry - IX

#### Fill In The Blanks



### → SOLUTION & SUSPENSION

- Solution is a homogenous mixture of two or more substances. 1.
- The most common solvent in nature is water. 2.
- An aqueous solution is the solution when the liquid solvent is water. 3.
- 10% M solution contains 10 gram solute, dissolved in 90 g solvent.
- M is the symbol for the concentration unit of molarity. 5.
- A solution is a homogenous mixture of two or more substances. 6.
- The component of solution present in smaller amount is called solute. 7.
- The component present in greater quantity is called solvent. 8.
- 9. When water is solvent solution is called aqueous solution.
- The states of matter solid, liquid and gas produce <u>nine</u> types of solutions on mixing. 10.
- Soft drinks are the examples of gas in liquid. 11.
- 12. Cloud and steam arc e examples of liquid in gas.
- 13. Smoke is solid in gas.
- 14. Milk in tea is liquid in liquid.
- The solubility of sugar in water at 0°C is 179 g/100 ml. 15.
- The solubility of gases in a liquid decreases with the inverse in temperature. 16.
- 17. The solubility of solids and liquids arc not affected by **pressure**.
- 18. The solubility of a gas in a liquid is directly proportional to the pressure gas. This is called Henry's Law.

- An unsaturated solution is that one which contains <u>less solute</u> than it has the capacity to <u>dissolve</u>.
- A <u>saturated</u> solution is that which contains <u>maximum</u> amount of solute in a given solvent at room temperature.
- The amount in grams of a solute required to be dissolved in 100 grams or 100 ml of a solvent to prepare a saturated solution at room temperature is called <u>solubility</u> of that solute.
- A super saturated solution is that which contains greater amount of dissolved solute than that present in a saturated solution t the given temperature.
- The process in which dissolved sol4te comes out of the solution and forms crystals is called <u>crystallization</u>.
- 24. Crystals are the homogenous solids, having regular and definite geometrical shape.
- For crystallization super <u>saturated solution</u> is prepared.
- Crystals of copper sulphate are called <u>blue vitriol</u>.
- 27. The formula for copper sulphate crystals is CuSO<sub>4</sub>,5H<sub>2</sub>O.
- In a solution the amount of solute dissolved in a given quantity of solvent is known as concentration.
- Molarity, Molality, mole fraction, percentage and normality are the common methods of expressing the <u>strength</u> of a solution.
- Molarity is defined as the number of <u>mole</u> of solute dissolved in <u>1</u> litre (1dm³) of a solution. Molarity is denoted by <u>M</u>.
- The Molarity of any solution is found by dividing the number of moles of solute by the number of litres of solution.
- 32. 1 mole of NaOH (40g) dissolved in 1 litre of solution is said to be 1 mole (1M) solution.
- 33. Number of moles of solute dissolved per 1000 (1kg) of solvent is called molality.
- 34. Molality is denoted by m.
- 35. 1 mole of NaCO<sub>3</sub> dissolved in 100 grams of solvent is said to 1 molal (1m) solution.

- Mole fraction is defined as the ratio of moles of solute or solvent to the total moles of the solution.
- 37. The sum of mole fractions of all components of a solution must be equal to 1.
- In <u>suspension</u> the solute particles are not soluble in solvent.
- Suspension is <u>heterogeneous</u> mixture.
- 40. Smoke is a suspension of Carbon particles in a gas or air.
- 41. Mud is a Suspension of fine particles of solid in a small quantity of liquid.
- 42. Foam (froth) is a suspension of fine particles of a gas in a liquid.
- 43. Emulsion is a suspension of droplets of one liquid into another in which it is not soluble.

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