

Chemistry



9th Class

Sindh Board Notes

Chapter # 9

Acids Bases And Salts



Fill In Blanks

پنجاب، سندھ، بلوچستان، خیبر پختونخواہ، بورڈز کے تمام نوٹس سابقہ پیپرز، اس ویب سائٹ سے فری ڈاؤن لوڈ کریں۔

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→ ACIDS, BASES AND SALTS

1. The formula of baking soda is NaHCO₃.
2. The formula of Epsom salt is MgSO₄·7H₂O.
3. $K_2SO_4Al_2(SO_4)_3 \cdot 24H_2O$ is the formula of Potash alum.
4. Molarity is the most convenient way of expressing concentration.
5. The molarity of solution is denoted by M.
6. A solution whose strength is known is called standard solution.
7. If H^+ concentration of a solution is 1×10^{-14} M, the solution is very basic.
8. If the OH ion concentration of a solution is 1×10^{-10} M, solution is medium acidic.
9. Titration is the process by which we can determine the concentration of un-known solution with the help of standard solution.
10. The solution whose H^+ ion concentration is 1×10^{-4} M, then its pH is 4.
11. The solution whose pH is 6, then its H ion concentration is 1×10^{-6} .
12. The volume of a pipette is generally 10 ml or cm^3 .
13. Molarity is defined as number of moles per litre of solution.
14. Acids were first recognized in 5th century.
15. In Latin, acidus means sour.
16. All citrous fruits contain large amount of ascorbic acid.
17. Vitamin C is also known as ascorbic acid and its formula is C₆H₈O₆.
18. Apples contain maleic acid. The souring of milk produces lactic acid.

19. The extract of Vinegar is acetic acid.
20. The consumption of H_2SO_4 is an index to the state of civilization and prosperity of a country.
21. HCl composes about 0.4% of gastric juice of our stomach.
22. Lye is commercial NaOH used for cleaning sink drains.
23. Lime water is a solution of Ca(OH)_2 .
24. Milk of Magnesia Mg(OH)_2 is used as an antacid, laxative and an antidote.
25. Acids have a sour taste. Acids react with bases to form salts and water.
26. Bases have bitter taste. Bases have slippery touch.
27. Bases react with acid to form salts and water.
28. Arrhenius first describe the nature of acids in his theory of ionization.
29. Arrhenius theory describes acid is a substance that yields H^+ ions when dissolved in water.
30. According to Arrhenius theory bases are the substances that yield OH^- ions on ionization.
31. Bronsted Lowry bases are proton acceptors. Bronsted Lowry acids are proton donors.
32. In Arrhenius theory describes acid is a substance that yields H^+ ions when dissolved in water.
33. According to Arrhenius theory bases are the substances that yield OH^- ions on ionization.
34. Lewis acid is any species which can accept a pair of electrons.
35. Lewis base is any species which can donate a pair of electrons.
36. When acids and bases react together in right proportions a neutral substance is formed.
37. Neutralization is the process, when acids and bases are mixed to form neutral product.
38. Neutralization is an exothermic reaction.

39. The number of ionizable hydrogen atoms present in a molecule of an acid is called **Basicity** of the acid. CH_2OOH is **mono** protic acid. H_2SO_4 is a **diprotic** acid. PO_4 is a **tri**-basic acid.
40. Acids that contain two or more acidic hydrogen atoms are called **Poly** protic acids.
41. The number of ionisable OH^- ions in a molecule of base is called **acidity** of the base.
42. Base that contain two or more OH^- ion per molecule are called **Poly-acid** bases.
43. $\text{Ca}(\text{OH})_2$ and $\text{Ba}(\text{OH})_2$ are **di-acid** bases. NaOH and **KOH** are **mono-acid** bases.
44. Bases that produce **3** moles of OH^- ions per mole of base are called tri-acid bases.
45. **$\text{Al}(\text{OH})_3$** and $\text{Cr}(\text{OH})_3$ are tri-acid bases.
46. Normal salts are formed by the complete **neutralization** of an acid by a base.
47. K_2SO_4 and NaCl are **normal** salts. KHCO_3 and NaHSO_4 are **acidic** salts.
48. Acidic salts have replaceable **hydrogen** ion. Basic salts have replaceable OH^- groups.
49. Two specific salts crystallize together to form **double** salts.
50. Double salts have definite number of **water** molecules with them.
51. Chrome Alum is a **double** salt. $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ is **carnalite**.
52. Formula of Mohar's salt is **$\text{FeSO}_4(\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$** .
53. Formula of **Chrome** Alum is $\text{K}_2\text{SO}_4 \cdot \text{Cr}(\text{SO}_4)_3 \cdot 24\text{H}_2\text{O}$.
54. Potash Alum is a **double** salt.
55. Sodium carbonate is commercially prepared by **Solvay** Process.
56. Na_2CO_3 is commonly known as **Soda Ash**.
57. Sodium carbonate deca hydrate ($\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$) is commonly known as **washing soda**.
58. Na_2SiO_3 is known as **water glass**.
59. NaHCO_3 is commonly known as **baking soda**.

60. NaFICO_3 is used in the preparation of effervescent drinks and fruit salts. NaHCO_3 is used in preparation of baking powder.
61. NaHCO_3 is used as fire extinguisher.
62. Copper sulphate is commonly known as blue vitriol or blue stone.
63. Magnesium sulphate is also known as Epsom salt.
64. It is used in making fire proof fabrics.
65. The formula of potash alum is $\text{K}_2\text{SO}\cdot\text{Al}(\text{SO}_4)_3\cdot 24\text{H}_2\text{O}$.
66. Water acts as an acid and as a base.
67. In presence of an acid water acts as a base. In the presence of base water acts as an acid.
68. A substance that can behave as both an acid and a base is called amphoteric substance.
69. K_w is called the ion-product constant of water.
70. Negative log of the H^+ ion concentration or H_3O_4 ion concentration is the pH of the solution.
71. $\text{pH} = \log [\text{H}^+]$
72. Negative log of the OH^- ion concentration is the pOH of the solution.
73. $\text{pOH} = -\log[\text{OH}^-]$
74. The sum of pH and pOH is always equal to 14.
75. Substance with pH value 2, 3, or 4 are medium acidic.
76. Substance with pH value 13, 14, 15 are very basic.
77. Substances with pH value 8, 9, are slightly basic.
78. Substance with pH value 7 is neutral substance.
79. Substance with PH value 10, 11, 12 are medium alkaline.
80. If the pH of blood drops to 7, the patient may go into coma.

81. The pH of human **blood** is normally between 7.35 -7.45.
82. pH rises as high as 7.7 or 7.8 causes **diabetes**.
83. pH of lemon juice is **2.3**. pH of vinegar is **2.8**.
84. pH of tomato juice is **4.2** pH of human urine is **5.0-7.0**.
85. pH of cow's milk is **6.5** pH of saliva is **7.0**.
86. pH of Egg white is **7.8**. pH of water is **7**.
87. A solution of known molarity is known as **standard** solution.
88. **Titration** is the process by which the concentration of unknown solution can be determined.
89. Titration plays an important role in determining amounts of **solutes** in a solution.
90. Concentration of a solution is expressed by **molarity**.
91. **Molarity** is defined as the number of moles of solute dissolved per litre of solution.
92. The point at which the titration is completed is called **end point**.
93. The solution in the burette is called **Titre**.
94. The solution in the titration flask is called **Titrant**.
95. Acid base **indicator** is used to detect the end point.

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